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Laboratory manual for physical chemistry where T_0 is the freezing point of the solvent. If BC in Fig. 1st-1 is taken to be a straight line, the

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slope dp_A/dT may be replaced by CB/BD , i.e., by $(p_A - p_s)/T$, so that $p_A = p_s + \frac{H_p}{RT} \Delta T$ Eq. 1st.4 The slope of the vapor pressure curve of the solid solvent is also given by another form of the

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The dilution formula allows you to determine the molarity of a solution after it has been diluted. The formula is: $M_1 V_1 = M_2 V_2$ Where M_1 and V_1 are the molarity and volume before a dilution, and M_2 and V_2 are the molarity and volume after a dilution. Consider this example from the 2006 NECTA: "Hydrochloric acid solution made by diluting

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Laboratory Manual of Practical Inorganic II Chemistry. Figure 1.1 Packing models of (a) side-by-side layer, (b) hexagon layer, (c) simple cubic layer, (d) body centered cubic. Determine the maximum amount of balls that can touch another one ball (center- ball) in the same layer for each arrangement.

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