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ionic: $H^+ (aq) + Cl^-(aq) + K$
 $+ (aq) + OH^-(aq) ? K^+ (aq) +$

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 $\text{Cl}^-(\text{aq}) + \text{H}_2\text{O}(\text{l})$ net
ionic: $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$
3. $\text{H}_2\text{SO}_4(\text{aq}) +$
 $\text{Mg}(\text{OH})_2(\text{s}) \rightarrow \text{MgSO}_4(\text{aq}) +$
 $2\text{H}_2\text{O}(\text{l})$ ionic: $2\text{H}^+(\text{aq}) +$
 $\text{SO}_4^{2-}(\text{aq}) + \text{Mg}^{2+}(\text{aq}) +$
 $2\text{OH}^-(\text{aq}) \rightarrow \text{Mg}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) +$
 $2\text{H}_2\text{O}(\text{l})$ net

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ionic: $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$
4. $\text{NH}_4\text{OH}(\text{aq}) + \text{HCl}(\text{aq}) \rightarrow \text{NH}_4\text{Cl}(\text{aq}) + \text{H}_2\text{O}(\text{l})$
ionic: $\text{NH}_4^+(\text{aq}) + \text{OH}^-(\text{aq}) + \text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{NH}_4^+(\text{aq}) +$

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different chloride solutions
are mixed with each other:
Solution 1 is an aqueous
solution of 4.85 g aluminum
chloride in 150. ml water

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Answer Key 2 has a volume of 175 ml and contains 7.75 g of zinc chloride. You mix the two solutions together in a large beaker. Solved:
Chem 1A: Chemical Reactions
And

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Answer CHEM 1A Review:

Experimental Errors,

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Name _____ Date _____

5. Iron III nitrate and potassium sulfide I

6. strontium acetate and phosphoric acid

- 7. aluminum sulfite and ammonium sulfide

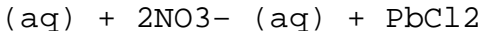
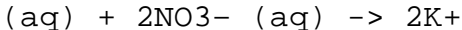
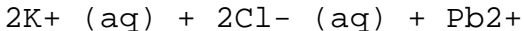
8. barium hydroxide and zinc(II)

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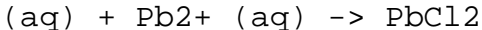
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(s) Net Ionic Equation: 2Cl^-



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(s) Directions: Write balanced molecular, ionic, and net ionic equations for each of the following reactions. Assume all reactions occur in aqueous solution.

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*Net Ionic Equation Worksheet
Answers*

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Answer Answer to 1a.

Identify whether each

species functions as a

Brønsted-Lowry acid or a

Brønsted-Lowry base in this

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net ionic equati... Solved:
1a. Identify Whether Each
Species Functions As A B ...
 $2\text{K}^+ (\text{aq}) + 2\text{Cl}^- (\text{aq}) + \text{Pb}^{2+}$
 $(\text{aq}) + 2\text{NO}_3^- (\text{aq}) \rightarrow 2\text{K}^+$
 $(\text{aq}) + 2\text{NO}_3^- (\text{aq}) + \text{PbCl}_2$
(s) Net Ionic

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barium sulfite (s) and
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